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Heisenberg uncertainty principle worksheet

Exercise 1: Heisenberg uncertainty principle [17 pts]

- a) Explain why the zero-point energy of the particle in a box is consistent with the uncertainty principle between position and momentum. [3 pts]
- b) Estimate the commutator $[\hat{p}, \hat{x}]$, where $\hat{p} = -i\hbar \frac{d}{dx}$ and $\hat{x} = x$. Choose (i) $\psi(x) = V$, a constant, (ii) $\psi(x) = \frac{1}{\sqrt{L}} e^{ikx}$. Considering Heisenberg's uncertainty principle, what do your results mean for the measurement of the energy and momentum of a particle in your system? [4 pts]
- c) An even function is a function that satisfies $f(-x) = f(x)$, an odd function satisfies $f(-x) = -f(x)$. Show that for the latter $\int_{-a}^a \psi(x) dx = 0$ (where a is a real positive number). Show that the product of two odd functions is an even function, the product of two even functions is an even function, and the product of an even and odd function is an odd function. [4 pts]
- d) A particle is in a state described by the normalized wave function

$$\psi(x) = \left(\frac{2}{\pi}\right)^{1/4} e^{-x^2} \quad (1)$$

where a is a constant and $-\infty < x < \infty$. Verify that the value of the product of errors $\Delta x \Delta p$ is consistent with the predictions from the uncertainty principle. Use the definition for the error (root-mean-square deviation) $\Delta p = \sqrt{\langle p^2 \rangle - \langle p \rangle^2}$ and $\Delta x = \sqrt{\langle x^2 \rangle - \langle x \rangle^2}$. Explain why this state is called a minimum uncertainty state. [6 pts]

Hint: Use your answer in (c) and some definite integrals given in the Math Sheet on Compendy.

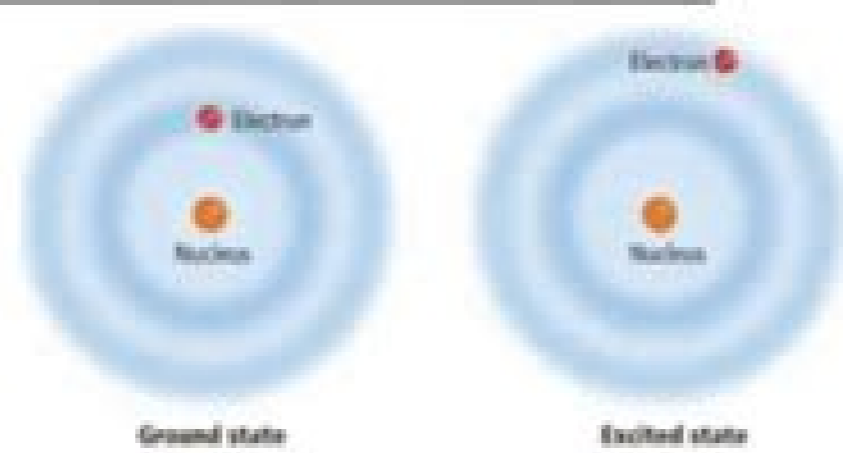
Section 2: Quantum Theory and the Atom - Notes

Objectives:

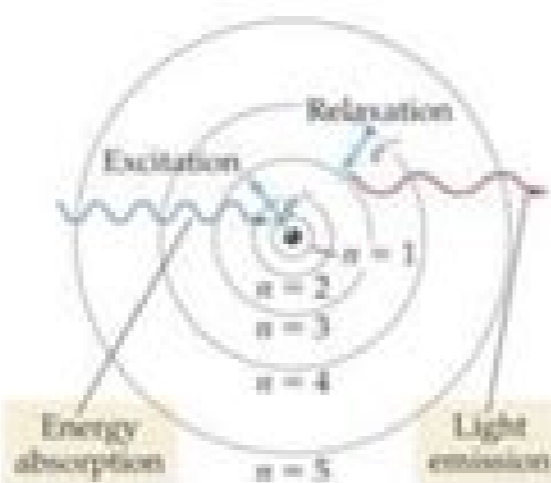
- Compare the Bohr and quantum mechanical models of the atom.
- Explain the impact of de Broglie's wave particle duality and the Heisenberg uncertainty principle on the current view of electrons in atoms.
- Identify the relationships among a hydrogen atom's energy levels, sublevels, and atomic orbitals.

Bohr's Model of the Atom:

- Einstein's theory of light's dual nature accounted for several unexplainable phenomena, but it did not explain why _____
- In 1913, _____, a Danish physicist working in Rutherford's laboratory, proposed a quantum model for the hydrogen atom that seemed to answer this question.
 - This model correctly predicted the frequency lines in _____ atomic emission spectrum.
 - The lowest allowable energy state of an atom is called its _____.
 - When an atom gains energy, it is in an _____.



- Bohr suggested that an electron moves around the nucleus only in certain allowed _____
- Each orbit was given a number, called the _____
 - Bohr orbits are like steps of a _____, each at a specific distance from the nucleus and each at a specific energy.



- Hydrogen's single electron is in the _____ orbit in the _____
 - When _____ is added, the electron moves to the _____ orbit.
- The electron releases _____ as it falls back towards the ground state.

Properties of EMR:

- All waves have distinct amplitudes, frequency, periods and wavelengths.
- All electromagnetic waves travel at the speed of light, $c = (3.0 \times 10^8 \text{ m/s})$.
- $c = \lambda \nu$
- The relationship between the energy and frequency is $E = h \nu$.
- All elements have their own individual atomic emission spectrum and absorption spectrum.

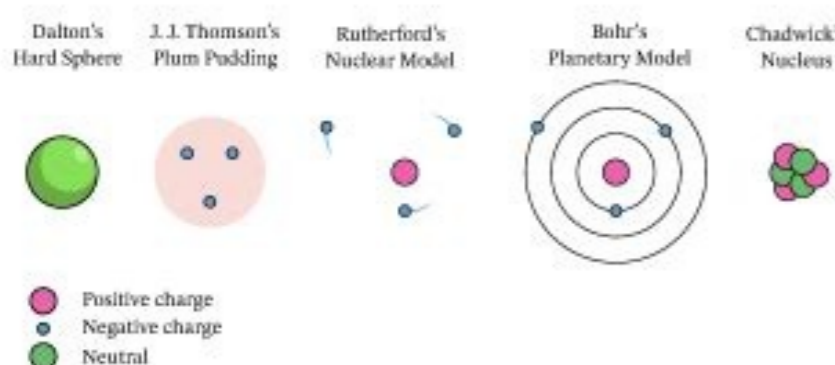


Explainer: Modern Atomic Theory

In this explainer, we will learn how to describe the concepts of modern atomic theory.

Each successive model of the atom was developed to produce a theoretical explanation for a practical result.

- ▶ John Dalton's hard-sphere model accounted for the way atoms pack together.
- ▶ J.J. Thomson's plum pudding model accounted for the presence of negatively charged particles in the atom, which we call "electrons."
- ▶ Geiger, Marsden, and Ernest Rutherford demonstrated the existence of the nucleus, a dense, positively charged part of the atom in the very center of it. It is the nucleus that contains protons.
- ▶ Niels Bohr and Ernest Rutherford proposed the Rutherford-Bohr model, often simply called the Bohr model, where electrons occupy orbits around the nucleus like the planets around the sun. This explained the features of the emission spectrum of hydrogen.
- ▶ James Chadwick later demonstrated that nuclei also contain uncharged particles. These are neutrons.



The Bohr model suggests the following about atoms:

- ▶ A dense positively charged nucleus sits in the middle of the atom.

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